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Combined Gas Laws $PV = \text{constant}$ $PV = \text{constant}$ (Boyles Law) T (combined gas law) $V = \text{constant}$ (Charles Law) $TP = \text{constant}$ TALL gas law problems must be work using absolute temperature, kelvin (K).Memory Work: 760 mmHg = 760 torr = 1 atm = 1.013 X 10⁵ Pa and $P_1V_1 = P_2V_2$ $T_1 T_2$ or $P_1V_1T_2 = P_2V_2T_1$ and STP (standard temperature and pressure) 0 oC and 1 atm (273 K) (760 torr)V Volume amount of gas cons...

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[CHAPTER 10](#)

Getting the other gas laws If temperature is constant If volume is constant you you get Boyles Law get Gay-Lussacs Law $P_1 V_1 = P_2 V_2$ $P_1 P_2 T_1 = T_2$ If pressure is constant you get Charless Law If all are variable, you get the Combined Gas Law. $V_1 V_2 T_1 = T_2 P_1V_1 P_2V_2 T_1 = T_2$ Boyles Law: Inverse relationship between Pressure and Volume. $P_1V_1 = P_2V_2$ P V Combined Gas Law: Gay-Lussacs Law: $P_1V_1 P_2V_2 \dots$

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