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Chemistry 31 Quiz 3 10 minutes March 20Name Lab Section NumberH2C2O4 Ka1 = 5.62 x 10<sup>2</sup>, Ka2 = 5.42 x 10<sup>5</sup> ; Kw = 1.00 x 10<sup>14</sup>. (5 pts) Calculate the pH of the following (Ignore activity and assume no secondary reactions besides acid base in water.): 0.015 M Na<sub>2</sub>C<sub>2</sub>O<sub>4</sub> [C<sub>2</sub>O<sub>4</sub><sup>2-</sup>] = 0.015 M See Chapter 6 problems 46, 48, and 6-1 1.00 x 10<sup>-14</sup> C<sub>2</sub>O<sub>4</sub><sup>2-</sup> + H<sub>2</sub>O ⇌ HC<sub>2</sub>O<sub>4</sub><sup>-</sup> + OH<sup>-</sup> Kb1 = 5 = 1.845 x 10<sup>10</sup> 5.42 x 10<sup>-10</sup> [HC<sub>2</sub>O<sub>4</sub><sup>-</sup>] [OH<sup>-</sup>]

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1) a.  $\log_2 32 = x$  b.  $\log_5 (1/5) = x$  c.  $\log_3 35 = 2x$  d.  $\ln x = -22$  a.  $2^3 = 8$  b.  $3^{-1} = 1/3$  c.  $e^x = 53$  a. 4 b. 5 c. -3 d. 3 e. 9  
 f. 34) a. 2 b. -2 c. 0 d. 4 e. 1 y) a. ? b. D: (-?, ?), R: (0, ?), HA @ y = 0 ? c. Inc: (-?, ?) d. f(x) shifted left 2, down 3 ? ? e. D: (-?, ?), R: (-3, ?), HA @ y = -3 x ? ? ? ? ? ? ? ? ? ? y ? ? ? ? ? ? 6) ? ? b. f(x): D: (0, ?), R: (-?, ?) g(x): D: (-?, 0), R: (-?, ?) ? g(x) ? ? f(x)...

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